

Engineering Thermodynamics Equation Sheet

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Engineering Thermodynamics Equation Sheet

Basic Thermodynamic Formulas (Exam Equation Sheet) Control Mass (no mass flow across system boundaries) Conservation of mass: $\dot{m}_1 = \dot{m}_2$. Conservation of energy (1st Law): $\dot{m}_1 h_1 + \dot{Q} = \dot{m}_2 h_2 + \dot{W}$

Basic Thermodynamic Formulas (Exam Equation Sheet)

File Type PDF Engineering Thermodynamics Formula Sheet Engineering Formula Sheet - Madison Local Schools For quasi-static and reversible processes, the first law of thermodynamics is: $dU = \delta Q - \delta W$ where δQ is the heat supplied to the system and δW is the work done by the system.

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ME 300: THERMODYNAMICS II BASIC EQUATION SHEET FALL 2020 Governing Principles Mass Balance: $\dot{m}_{in} - \dot{m}_{out} = \frac{dm_{cv}}{dt}$ Closed System: $m = \text{constant}$ Energy Balance: $\dot{Q} - \dot{W} = \dot{m} \left(h_2 - h_1 + \frac{V_2^2 - V_1^2}{2} + g(z_2 - z_1) \right)$

ME 300: THERMODYNAMICS II BASIC EQUATION SHEET FALL 2020 ...

ME 211 and ME312 Thermodynamics Equation Sheet D. Abata, April 1, 2020 Conservation of mass: where Boundary work any system: and flow work (open system) , assuming ideal gas and since $T=C$ then and For the polytropic process, that is : Open system work: , ,

ME 211 and ME312 Thermodynamics Equation Sheet

General equation . Valid at any instance of time Steady or not steady flow. Usually Simplifies to $\dot{Q} - \dot{W} = \dot{m} \left(h_2 - h_1 + \frac{V_2^2 - V_1^2}{2} + g(z_2 - z_1) \right)$ State 1 State 2 Second law Non- flow $\dot{m} s_2 - \dot{m} s_1 - \frac{\dot{Q}}{T_s} = \dot{m} s_2 - \dot{m} s_1 - \dot{q} T_s$ gen steady transient 0 $\dot{m} s_2 - \dot{m} s_1 - \dot{q} T_s = \dot{m} s_2 - \dot{m} s_1 - \dot{q} T_s$ gen

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This list gives you some of the most common conversion factors you need in thermodynamics. Acceleration: $1 \text{ m/s}^2 = 100 \text{ cm/s}^2$. Area: $1 \text{ m}^2 = 10^4 \text{ cm}^2 = 10^6 \text{ mm}^2$. Density: $1 \text{ g/cm}^3 = 1 \text{ kg/L} = 1,000 \text{ kg/m}^3$. Energy, heat, work, internal energy, enthalpy: $1 \text{ kJ} = 1,000 \text{ J} = 1,000 \text{ N}\cdot\text{m} = 1 \text{ kPa}\cdot\text{m}^3$. $1 \text{ kJ/kg} = 1,000 \text{ m}^2/\text{s}^2$.

Thermodynamics For Dummies Cheat Sheet - dummies

Thermodynamics $\dot{Q} = \frac{\Delta T}{R}$ $\dot{Q} = h A \Delta T$ $\dot{Q} = \sigma A T^4$ $\dot{Q} = \frac{A v P}{R}$ = rate of heat transfer Q = thermal energy A = Area of thermal conductivity U = coefficient of heat conductivity (U-factor) T = change in temperature R = resistance to heat flow (R-value) v = velocity net = net power radiated = $5.6696 \times 10^{-8} \text{ W/m}^2 \cdot \text{K}^{-4}$

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temperature at time 1, time 2 $v =$ flow velocity

Engineering Formula Sheet

Purdue's School of Mechanical Engineering conducts world-class research in robotics, automotive, manufacturing, rocket and jet propulsion, nanotechnology, and much more. ME 200 - Thermodynamics I - Purdue University Mechanical Engineering

Thermodynamics I - College of Engineering - Purdue University

$S = k_B (\ln \Omega)$, where k_B is the Boltzmann constant, and Ω denotes the volume of macrostate in the phase space or otherwise called thermodynamic probability. $dS = \frac{\delta Q}{T}$, for reversible processes only.

Table of thermodynamic equations - Wikipedia

All of the thermodynamics in one sheet Figure 1: thermodynamics. Figure 2: polytropic process diagrams. Figure 3: first and second laws diagrams.

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or Pa) Absolute Temperature, T ($^{\circ}R$ or K) Volume, V (ft^3 or m^3) v (ft^3/lbm or m^3/kg) Internal Energy, U (Btu or kJ) u (Btu/lbm or kJ/kg) Enthalpy, H (Btu or KJ) Enthalpy, $h = u + Pv = H/m$ (usually in Btu/lbm or kJ/kg) Entropy, S ($Btu/^{\circ}R$ or kJ/K)

FE Reference 8-2.1104web - College of Engineering

Thermodynamics key facts (7/9) • Ideal gas law • 1. st. form : $PV = nRT$ • P = Pressure, V = Volume, n = number of moles, R = Boltzmann's constant, T = temperature [in K] • 2. nd. form : $PV = Nk_B T$ • N = number of molecules, k_B = Boltzmann's constant

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Revision : Thermodynamics

For quasi-static processes where changes in kinetic and potential energy are not important. $dU = Q - pdV$ or $du = q - pdv$ $dH = Q + Vdp$ or $dh = q + vdp$

COMPENDIUM OF EQUATIONS Unified Engineering Thermodynamics

Thermodynamics is filled with equations and formulas. Here's a list of the most important ones you need to do the calculations necessary for solving thermodynamics problems. Combustion equations: Air-fuel ratio: Hydrocarbon fuel combustion reaction: Compressibility calculations: Compressibility factor Z: $Pv = ZRT$ Reduced temperature: Reduced pressure: Pseudo-reduced specific volume ...

Important Thermodynamic Equations and Formulas - dummies

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Thermodynamics Equations Cheat Sheet

MEASURED THERMODYNAMIC PROPERTIES AND OTHER BASIC CONCEPTS | 5 1. MEASURED THERMODYNAMIC PROPERTIES AND OTHER BASIC CONCEPTS 1.1 PRELIMINARY CONCEPTS - THE LANGUAGE OF THERMODYNAMICS In order to accurately and precisely discuss various aspects of thermodynamics, it is essential to have a well-defined vernacular. As such, a list of some foundational concepts and their definitions are shown

Chemical Engineering Thermodynamics

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