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answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of H

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ANSWER KEY 4 49.3%  
Rh 23.4 % C 27.3 % N  
5. a) PBr 5 b) Zr(BO 3)  
2 Zirconium (VI) borate  
6. a) C 3H 5Cl b) C 6H  
10Cl 2 7. empirical:  
KCO 2 molecular: K 2C  
2O 4 name: Potassium  
oxalate 8. You will be  
setting up a balanced  
equation. Using the  
Law of Conservation of  
mass (FIRST!) you will  
be using Mole

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stoichiometry to  
determine the  
balanced equation.

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